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Chapter 6 Study Guide Chemistry - Lesson Worksheets

Accelerated Chemistry Anderson - MCHS Modern Chemistry 1 Chemical Bonding CHAPTER 6 Chemical Bonding SECTION 1 Introduction to Chemical Bonding OBJECTIVES 1. Define Chemical bond. 2. Explain why most atoms form chemical bonds. 3. Describe ionic and covalent bonding.. 4.

CHAPTER 6 Chemical Bonding

CHAPTER 6 REVIEW Chemical Bonding SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. a The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2. d In a crystal of an ionic compound, each cation is surrounded by a number of (a) molecules. (c) dipoles. (b) positive ions. (d) negative ions.

6 Chemical Bonding

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Chapter 6 Chemistry Review Answer Key

Vocabulary from Chapter 6 of the Modern Chemistry textbook. (30 total words) Terms in this set (30) chemical bond. the attractive force that holds atoms or ions together. ionic bonding. a force that attracts electrons from one atom to another, which transforms a neutral atom into an ion.

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Modern Chemistry Chapter 6 Section 2. STUDY. PLAY. single bond. is a covalent bond in which one pair of electrons is shared between two atoms. multiple bond. double and triple bonds are referred. molecule. a neutral group of atoms that are held together by covalent bonds. molecular compound.

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Ch 6 - HonorsChemWins Section 6.1 - Ionic Bonds When the highest occupied energy level of an atom is filled with electrons, the atom is stable and not likely to react. The noble gases have stable electron configurations. Read Free Chapter 6 Chemical Bonding Section 4 Worksheet Answers.

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Group 2 elements in the modern periodic table and are highly reactive. Transition Metals An element in groups 3-12 that is contained in the d-block of the periodic table and, with some exceptions, is characterized by a filled outermost s orbital energy level n, and filled or partially filled d orbitals of energy level n -1.

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