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ANSWER Answer the

following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

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SECTION 3 Date
CHAPTER 11 REVIEW
Gases Class SHORT
ANSWER Answer the
following questions in
the space provided. c c
The molar mass of a
gas at STP is the
density of that gas (a)
multiplied by the mass

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of 1 mol. (b) divided by

the mass of 1 mol. nRT

(c) multiplied by 22.4

L. (d) divided by 22.4 L.

For the expression $V =$

(a) increasing P

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...

3 react according to the above equation, how many moles of NO will be consumed? 8.

Propyne gas can be used as a fuel. The combustion reaction of propyne can be represented by the following equation:

$$\text{C}_3\text{H}_4(\text{g}) + 4\text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$$

a. Write

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all the possible mole
ratios in this system. 4
mol O 2:1 mol C 3H 4;
3 mol CO 2:1 mol C 3H
4; 2 ...

mc06se cFMsr i-vi

Other Results for
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REVIEW The Periodic
Law SECTION 1 SHORT
ANSWER Answer the

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following questions in the space provided. 1.
c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass.

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CHAPTER 3 REVIEW

Atoms: The Building
Blocks of Matter

SECTION 3 SHORT

ANSWER Answer the

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following questions in
the space provided. 1.

Explain the difference
between the mass
number and the atomic
number of a nuclide.

Mass number is the
total number of
protons and neutrons
in the nucleus of an
isotope.

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SECTION 3 continued

5. Explain the following

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properties of solids by describing what is occurring at the atomic level. a. Metallic solids conduct electricity well, but covalent network solids do not. Metals have many electrons that are not bound to any one atom; therefore they are able to move throughout the crystal.

**10 States of Matter -
Ms. Agostine's
Chemistry Page**

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Name Section Quiz,

continued Class Date

5. In some instances, the concentration of a solution is expressed as molality instead of molarity because a. molality is easier to calculate. b. molarity applies only to solid-liquid solutions.

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Chemical Bonding

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CHAPTER 6 STUDY

GUIDE Chemical

Bonding SECTION 3

IONIC BONDING AND

IONIC COMPOUNDS

SHORT ANSWER

Answer the following questions in the space provided. 1. _____ The notation for sodium chloride, NaCl, stands for one (a) formula unit.

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CHAPTER 3 REVIEW

Atoms: The Building

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SECTION 3 SHORT

ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

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