

2 Chapter 9 Stoichiometry Answers Section 2

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CHAPTER 9 REVIEW Stoichiometry
MIXED REVIEW SHORT ANSWER Answer
the following questions in the space
provided. 1. Given the following
equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ 4 a. What is the value of

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2 the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of C₃H₄? 2 mol O₂:1 mol H₂O c. What is the mole ratio of O₂ to H

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Chapter 9 - Stoichiometry 9-1

Introduction to Stoichiometry

Composition Stoichiometry - deals with mass relationships of elements in compounds
Reaction Stoichiometry -

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2 Involves mass relationships between reactants and products in a chemical reaction I. Reaction Stoichiometry Problems A.

Chapter 9 - Stoichiometry

Modern Chemistry 77 Stoichiometry
CHAPTER 9 REVIEW Stoichiometry
SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. _____ The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H

CHAPTER 9 REVIEW Stoichiometry

Chapter 9 Mixed Review Stoichiometry
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SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g

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Introduction to Stoichiometry SECTION 2
Ideal Stoichiometric Calculations
SECTION 3 Limiting Reactants and
Percentage Yield Why It Matters Video ...
The number of significant figures in the
answer Chapter 9 DO NOT
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CHAPTER 9 REVIEW. Stoichiometry. SECTION 9.2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many grams of O_2 form if 3.0 mol of KClO_3 are totally consumed? 2. Given the following equation ...

CHAPTER 9 REVIEW

Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of

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2 reactants/products involved in a chemical reaction. Students...

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Stoichiometry. SECTION 2. PROBLEMS

Write the answer on the line to the left.

Show all your work in the space

provided. 1. The following equation represents a laboratory preparation for oxygen gas:



How many moles of O_2 form if 3.0 mol of KClO_3 are totally consumed?

2. Given the following equation: $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \rightarrow 2\text{HF}(\text{g})$

CHAPTER 9 REVIEW

CHAPTER 9 REVIEW. Stoichiometry.

SECTION 1. SHORT ANSWER Answer the

following questions in the space

provided. 1. _____ The coefficients in a chemical equation represent the.

(a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each

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2 compound in a reaction.

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SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided
1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield
2 60 mol of N_2 are mixed with 120 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g)$

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2 SECTION 2 continued Date Class _____
60.2 g of lithium chloride reacts with 100. g of oxygen gas to produce lithium chloride and oxygen gas. The oxygen gas produced in part a has a density of 1.43 g/L. Calculate the volume of this gas.
76 STOICHIOMETRY MODERN CHEMISTRY a.
— 81 g of nitrogen gas ...

Date. FCHAPJ REVIEW.

CHAPTER 9 REVIEW Stoichiometry
SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided
1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield
2 60 mol of N_2 are mixed with 120 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g)$

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Problem Type 4: Given is a mass and
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substance (g) Problem Type 3: Given is a
mass and unknown is an amount in
moles Mass of given substance (g)
Reaction Stoichiometry Problems,
continued Section ...

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