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The color becomes darker as N 2 O 4 is converted to NO 2.When the system reaches equilibrium, both N 2 O 4 and NO 2 are present (). Figure 1. A mixture of NO 2 and N 2 O 4 moves toward equilibrium. Colorless N 2 O 4 reacts to form brown NO 2.As the reaction proceeds toward equilibrium, the color of the mixture darkens due to the increasing concentration of NO 2.. The formation of NO 2 from N ...

13.1 Chemical Equilibria - Chemistry
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Equation 13.5.6 is the same as the equation for the reaction rate of a first-order reaction (Equation 13.3.2), except that it uses numbers of atoms instead of concentrations. In fact, radioactive decay is a first-order process and can be described in terms of either the differential rate law (Equation 13.5.7) or the integrated rate law: